

Negatively charged silver nanoparticles with potent antibacterial activity and reduced toxicity for pharmaceutical preparations

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Background: The discovery of new solutions with antibacterial activity as efficient and safe alternatives to common preservatives (such as parabens) and to combat emerging infections and drug-resistant bacterial pathogens is highly expected in cosmetics and pharmaceuticals. Colloidal silver nanoparticles (NPs) are attracting interest as novel effective antimicrobial agents for the prevention of several infectious diseases.

Methods: Water-soluble, negatively charged silver nanoparticles (AgNPs) were synthesized by reduction with citric and tannic acid and characterized by transmission electron microscopy, dynamic light scattering, zeta potential, differential centrifuge sedimentation, and ultraviolet-visible spectroscopy. AgNPs were tested with model Gram-negative and Gram-positive bacteria in comparison to two different kinds of commercially available AgNPs.

Results: In this work, AgNPs with higher antibacterial activity compared to the commercially available colloidal silver solutions were prepared and investigated. Bacteria were plated and the antibacterial activity was tested at the same concentration of silver ions in all samples. The AgNPs did not show any significant reduction in the antibacterial activity for an acceptable time period. In addition, AgNPs were transferred to organic phase and retained their antibacterial efficacy in both aqueous and nonaqueous media and exhibited no toxicity in eukaryotic cells.

Conclusion: We developed AgNPs with a 20 nm diameter and negative zeta potential with powerful antibacterial activity and low toxicity compared to currently available colloidal silver, suitable for cosmetic preservatives and pharmaceutical preparations administrable to humans and/or animals as needed.

Keywords: silver nanoparticles, antibacterial activity, long-term effect, nanoparticle toxicity, phase transfer

Introduction

The discovery and introduction of antibiotics was one of the greatest conquests in the history of medicine considerably contributing to improve the average life expectancy of the world population.¹ However, the increasing misuse and abuse of antibacterial agents has led to resistance and cross-resistance. Indeed, most pathogenic bacteria become resistant with time to the action of antibiotics.^{2,3} This effect represents a paramount medical problem so that new molecules and compounds with high and persistent antibacterial activity are required to face the challenge of emerging infections and the global spread of drug-resistant microbial pathogens.⁴

In addition to the treatment of infections, novel preservatives are essential in pharmaceutical and cosmetic areas as an alternative solution to the widely used esters of *p*-hydroxybenzoic acid (parabens).⁵ Despite their potent antimicrobial action, parabens

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are known to bind to human estrogen receptors, although with lower affinity compared to estradiol, potentially leading to undesired side effects, including acting as promoters of breast cancer.^{6,7} For these reasons, the use of nontraditional antibacterial compounds is assuming increasing relevance.⁸ Taking advantage of the bacteria sensitivity to silver ions,⁹ colloidal silver nanoparticles (CSNPs) have emerged as a novel class of potential antimicrobial agents for the treatment of different infectious diseases, including antibiotic-resistant pathogens, and have been assessed *in vitro* as well as *in animal*.¹⁰

Due to its unique chemical properties, silver is broadly utilized in several different fields, for example, optics and photonics, electron microscopy, catalysis, and biology; silver is found in everyday products such as computers, cell phones, vacuum cleaners, filters for water and air, textile products, and due to the antibacterial and antimicrobial properties, also in deodorants, soaps, toothbrushes, and toothpaste.¹¹ Silver, in the form of silver ions or nanoparticles (NPs), has been used for centuries against various diseases and, in particular, for its strong antiseptic and antimicrobial efficacy against both Gram-positive and Gram-negative bacteria associated with low toxicity. Although the antimicrobial mechanism of action of silver ions has not been fully elucidated yet and remains controversial, evidence of their activity in the inhibition of the enzyme respiratory system, in the alteration of the microbial DNA, and in the cell wall degradation has been provided.¹² Moreover, silver nanoparticles (AgNPs) exhibit a broad spectrum of actions against morphologically and metabolically different microorganisms.¹³ The differences in using AgNPs and silver ions as salts are essentially associated to a unique character of the NPs in biological systems, which allows us to assume the occurrence of effective interactions with bacterial membranes and sustained intracellular release of silver ions. In particular, the chemical–physical properties of AgNPs allow for their interaction with the surface of the bacterial cell by altering the structure therein, thus favoring its permeability and subsequent cell death.^{4,13}

To develop AgNPs as additives suitable to be included in the production of cosmetics or pharmaceutical formulations for the purpose of human care, we have selected a chemical synthetic approach for their preparation in different suspensions, including aqueous and oily solutions. In this article, we report the synthesis, characterization, cytotoxicity, and potent antibacterial activity of 20 nm, negatively charged AgNPs in comparison with two different colloidal silver formulations from commercial sources.

Materials and methods

Materials

Silver nitrate ($\geq 99.0\%$), tannic acid, sodium citrate (SC), octadecylamine, silver dispersion (SNPs, 40 nm particle size), and all organic solvents were purchased from Sigma-Aldrich (St Louis, MO, USA). All chemicals were used as received without further purification. Paraffin was obtained from Marco Viti (Mozzate, Italy). Colloidal NPs (CNPs) were purchased from Atena Srl (Toronto, Ontario, CA).

AgNPs synthesis

An aqueous solution (100 mL) containing SC (5 mM) and tannic acid (0.025 mM) was prepared and heated to the boiling point. The solution was boiled for 15 min, and then 1 mL of aqueous silver nitrate solution (25 mM) was quickly added with a syringe under vigorous stirring (750 rpm). The solution became immediately yellow, and after 1 h heating, it was allowed to cool down and kept under stirring overnight at room temperature.

AgNPs formation was assessed using ultraviolet–visible (UV–vis) spectroscopy by the occurrence of the characteristic absorption peak corresponding to the plasmon resonance band. The resulting AgNPs were purified by centrifugation (two washes: 25,000× *g*, 45 min, 4°C) and further suspended in 1 mM SC solution or 0.3 mM phosphate buffer (PB), pH 7.2, depending on the use.

NPs characterization

Dynamic light scattering and zeta potential

The NPs hydrodynamic diameter and zeta potential were analyzed on a Zetasizer Nano ZS ZEN3600 from Malvern Instruments Ltd (Worcestershire, UK) operating at a light source wavelength of 633 nm and a fixed scattering angle of 173°. For dynamic light scattering (DLS) analysis, the purified samples were diluted three times both in SC 1 mM and in 0.3 mM PB, pH 7.2. For zeta potential, the samples were diluted in 0.3 mM PB, pH 7.2. The results were expressed as mean \pm standard deviation (SD) of three measurements.

Transmission electron microscopy

NPs were visualized using an FEI 120 KV Tecnai G2 spirit Bio TWIN Instrument. The sample (2 μ L) was deposited onto a formvar-coated 200-mesh copper grid (Ted Pella, Redding, CA, USA), and allowed to dry in air before examination. For the mean diameter determination of NPs, images were processed with ImageJ software. The reported value was calculated measuring at least 100 particles.

UV–vis spectroscopy

UV–vis spectra were recorded using a Nanodrop 2000c spectrophotometer from Thermo Scientific (Wilmington, NC, USA). Before measurement, all samples were diluted three times in 1 mM SC and in 0.3 mM PB, pH 7.2.

Differential centrifuge sedimentation analysis

Particle size distribution was measured using a CPS disc centrifuge DC24000 (CPS Instruments Inc.). A gradient fluid, 8–24 wt% sucrose solution in Milli-Q water, was freshly prepared and filled successively in nine steps into the disc, rotating at a speed of 24,000 rpm, starting with the solution of the highest density. Calibration was performed using poly(vinyl chloride) particles (0.239 μm , Analytik, Ltd) as a calibration standard before each measurement.

Silver content quantification

The total silver content in each kind of NP sample was determined by inductively coupled plasma optical emission spectrometry (ICP-OES). For this purpose, 3 mL of *aqua regia* were added to 200 μL of each sample and, after 72 h digestion, the samples were diluted with 7 mL of distilled water. All samples were measured in triplicate with the Optima 7000 DV ICP-OES.

Stability test

AgNPs stability (4°C, 6 months) was assessed in 1 mM SC and in 0.3 mM PB, pH 7.2, using UV–vis spectroscopy. Absorption spectra were detected and reported in the supporting information along with maximum absorbance as a function of time.

Lyophilization

NPs suspension in 1 mM SC was dried by lyophilization using a Christ Alpha 1–2 D freeze dryer from Martin Christ Gefriertrocknungsanlagen GmbH (Osterode an Harz, Germany). NPs were characterized as described above.

Antibacterial activities test

Antibacterial activity of NPs suspensions was tested on Gram-negative (*Escherichia coli* MG1665) and Gram-positive (*Staphylococcus aureus*) bacteria. A typical procedure was as follows: bacterial cultures were grown overnight into an Erlenmeyer flask containing nutrient broth (Bactotryptone 10% w/V, yeast extract 5%, NaCl 5%), then the suspension was diluted at an initial optical density of 0.05 at 600 nm ($\text{OD}_{600\text{ nm}}$) and cultures were allowed to grow at 37°C under

stirring (160 rpm). When the cultures reached an $\text{OD}_{600\text{ nm}}$ of 0.65 ($\sim 10^8$ colony forming units (CFU)/mL), 1 mL was withdrawn, centrifuged (2,000 $\times g$, 15 min, 25°C), and washed twice with 0.3 mM PB, pH 7.2. Following washing, the cells were suspended and diluted in PB containing AgNPs or other antibacterial agents to get a final bacterial concentration of 2×10^5 CFU/mL. Untreated samples, diluted in PB, were used as control. Bacterial suspensions were incubated on a rotary shaker at 130 rpm, 37°C for 2 h. Thereafter, 100 μL of decimal dilution (10- and 10²-fold) were plated onto Nutrient Agar plates. Following incubation (12 h at 37°C), the plates were counted for viable bacteria. The survival rate was calculated according to the equation:

$$\text{Fold-logarithmic reduction} = \log N_{\text{untreated}} / N_{\text{treated}}$$

where $N_{\text{untreated}}$ is the number of CFU/mL in the control and N_{treated} is the number of CFU/mL in the samples following exposure to an antibacterial agent. Each experiment was performed in triplicates.

To compare their efficacy, different antibacterial agents were tested at the same silver content: 0.54 ng/mL for *E. coli* and 5.4 ng/mL for *S. aureus*, respectively.

Long-term antibacterial activity

The experiment was performed as describe above. In this case, 2×10^5 cells were incubated with antibacterial agents (0.54 ng/mL silver content) for 2, 4, 6, and 24 h. After the incubation time, 2×10^2 theoretical cells number of each samples were plated on a solid growth medium and were incubated overnight at 37°C in duplicate. Grown CFU were counted and the results expressed as follows:

$$\text{Inhibition effect (\%)} = N_{\text{untreated}} / N_{\text{treated}} \times 100$$

The antibacterial activity was evaluated only for *E. coli* because *S. aureus* did not show survival at long incubation time.

Phase transfer

Octadecylamine solution in ethanol (10 mL, 20 mM) was added while stirring (500 rpm) to 10 mL of unpurified AgNPs solution obtained as described above. After 5 min, toluene (10 mL) was added to the mixture and left under stirring for 10 minutes. The AgNPs were transferred into the organic phase. No NPs were then recovered both in the aqueous phase and at the interphase. The solvent was evaporated and the product was dispersed in

ethanol. To remove the excess alkylamine, the resultant AgNPs were purified by centrifugation (two washes, $3,500\times g$, 15 min, at 4°C). After solvent evaporation, the AgNPs were suspended in tetrahydrofuran (THF) or an oily solution (eg, peanut oil) depending on the use.

Scanning and transmission electron microscopy analyses

Bacterial suspension was prepared as described in the anti-bacterial activity test section. When the suspension reached an $\text{OD}_{600\text{ nm}}$ of 0.65, 15 mL were withdrawn, centrifuged ($2,000\times g$, 15 min, 25°C), and washed twice with 0.3 mM PB, pH 7.2. Following washing, 10^8 cells were resuspended in 30-mL buffer containing (treated) or not containing (untreated) AgNPs at a final concentration of $0.54\ \mu\text{g/mL}$ for *E. coli* and $5.4\ \mu\text{g/mL}$ for *S. aureus*. The bacterial suspension was incubated on a rotary shaker at 130 rpm, 37°C , for 2 h. After incubation, the cells were washed three times by centrifugation ($2,000\times g$, 7 min) to separate the residual AgNPs. Next, each sample (both pellets and bacterial suspensions recovered after treatment) was dispersed in 1 mL of 2.5% w/V glutaraldehyde solution.

Transmission electron microscopy (TEM) samples were prepared as follows: the suspension was washed three times in PB and resuspended in double-distilled water. The sample (2 μL) was deposited onto a formvar-coated 200-mesh copper grid (Ted Pella, Redding, CA, USA) and allowed to dry in air before examination. In case of cell pellets, after one rinsing with PB, the specimens were postfixed in 1.5% osmium tetroxide for 2 h, dehydrated by 70%, 90%, and 100% ethanol, and embedded in epoxy resin (PolyBed 812 Polysciences Inc., Warminster, PA, USA). The ultrathin sections were stained with uranyl acetate and lead citrate and examined.

For SEM analysis, after washing, the bacterial suspensions were dehydrated by 70%, 90%, and 100% ethanol and resuspended in hexamethyldisilazane. Several drops of the suspension were deposited on a glass slide, dried, and sputter-coated in a vacuum with an electrically conductive 5 nm thick layer of gold–palladium alloy. SEM images were then recorded with a SEM (Leica S420).

MTT assay

3T3-L1 murine fibroblasts were seeded on a 96-multiwell dish at a density of 5,000 cells/ cm^2 and grown for 24 h in Dulbecco's Modified Eagle's Medium (DMEM, supplemented with 10% fetal bovine serum, 2 mM L-glutamine, 50 UI/mL penicillin, and 50 mg/mL streptomycin). The cells were incubated with AgNPs at different concentrations (1 and

10 ng/mL). At the indicated time points (24, 48, and 72 h), the cells were washed with PBS and then incubated for 3 h at 37°C with 0.1 mL of 3-(4,5-dimethyl-2-thiazolyl)-2,5-diphenyl-2H-tetrazolium bromide (MTT) stock solution previously diluted 1:10 in DMEM without phenol red. At the end of the incubation, 0.1 mL of MTT solubilizing solution was added to each well to solubilize the MTT formazan crystals (Sigma-Aldrich). Absorbances were read immediately in an UV–vis plate reader using a test wavelength of 570 nm and a reference wavelength of 620 nm. The results were expressed as mean \pm standard deviation of an average of four individual experiments. The statistical significance was determined using one-way analysis of variance (ANOVA).

Results and discussion

Synthesis and characterization of AgNPs

Long-term stable aqueous colloidal dispersions of AgNPs with narrow size distributions were synthesized according to Bastús et al, with some modifications.¹⁴ An aqueous silver nitrate solution was boiled in a reducing agent mixture containing SC and tannic acid. Upon silver addition, the solution immediately turned clear yellow due to the CNPs formation (Figure 1A).

The reaction was monitored by UV–vis spectroscopy. A peak ~ 400 nm corresponding to the Ag surface plasmon resonance appeared associated to the formation of NPs (Figure 1A). The thin shape (full width at half maximum = 68 nm) and peak symmetry reflected the AgNPs narrow distribution in terms of size and shape. The observed absorption λ_{max} at 405 nm was in line with the NPs size expected from the results of the relevant TEM and DLS analyses.^{15,16} In addition, a great reproducibility intra-batches was observed confirming the method robustness (Table S1; Figures S1 and S2).

The size and net particle charge of the as-synthesized AgNPs were determined through different techniques. The AgNPs showed a mean effective diameter of 19.3 ± 5.6 nm assessed by TEM (Figure 1C); a mean hydrodynamic size of 42.3 ± 4.7 nm (PDI = 0.204) evaluated by DLS (Figure 1B); and a centrifugal density size of 16.6 ± 3.7 nm obtained by differential centrifugal sedimentation (DCS; Figure 1D). Although the size values derived from TEM and DCS were comparable, the significant difference observed by DLS might be due to a strong solvation effect. As expected, due to the presence of citrate as stabilizer, the AgNPs were negatively charged with a zeta potential of -18.3 ± 1.9 mV calculated by electrophoretic mobility. The long-term stability of AgNPs was assessed by monitoring the NPs plasmon band position of AgNPs by UV–vis absorption spectroscopy. Either 1 mM SC solution

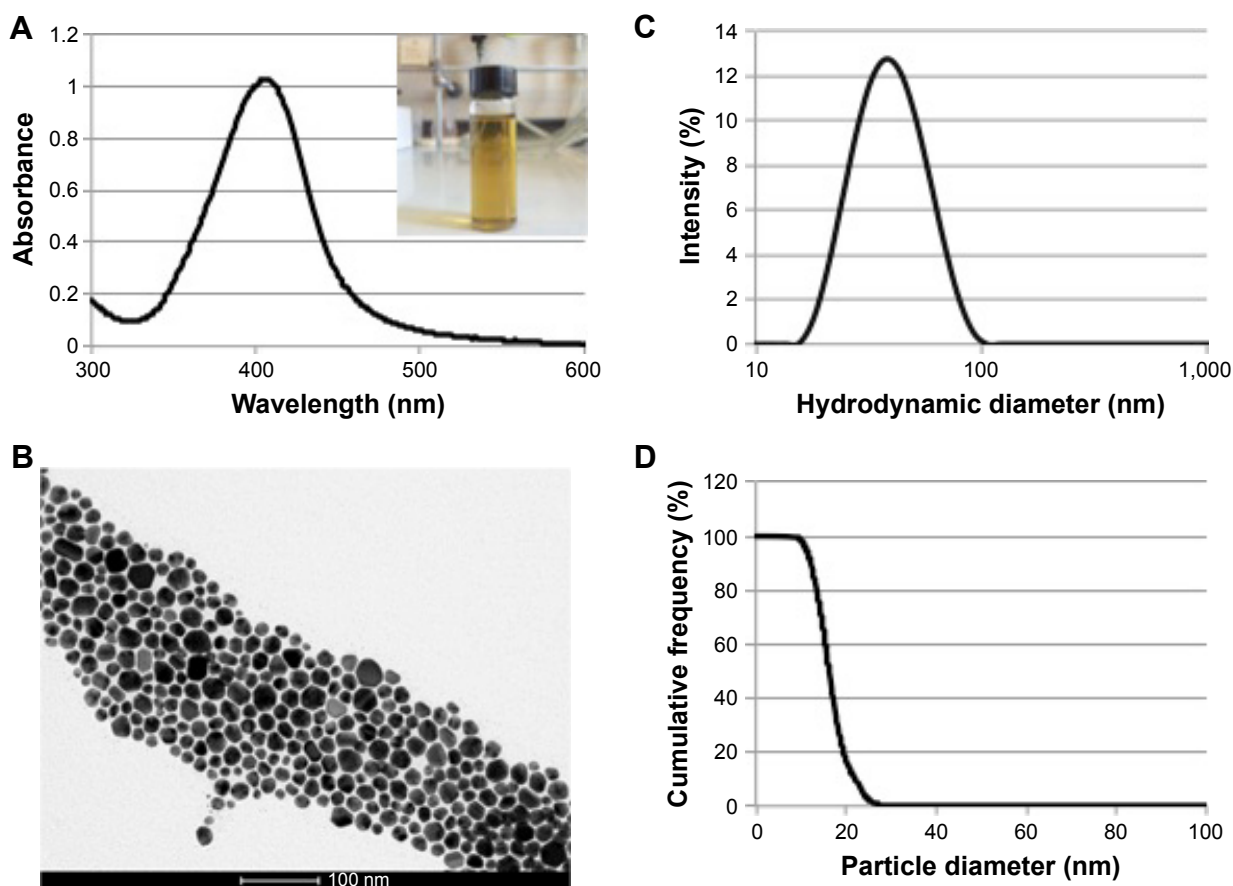


Figure 1 AgNPs characterization.

Notes: (A) Plasmon band position of AgNPs by UV-vis absorption spectroscopy and AgNPs colloidal solution (inset); (B) TEM image of AgNPs suspended in water, scale bar: 100 nm. AgNPs size distribution evaluated by (C) DLS in 1 mM citrate and (D) DCS in a sucrose gradient.

Abbreviations: AgNPs, silver nanoparticles; DCS, differential centrifugal sedimentation; DLS, dynamic light scattering; TEM, transmission electron microscopy; UV-vis, ultraviolet-visible.

or 0.3 mM PB, pH 7.2 was used to assess the reliability of the antibacterial test conditions. No degradation or aggregation was observed within 6 months of storage (Figure S3). Importantly, in view of a possible industrial production and purification, the integrity and colloidal stability of AgNPs was confirmed by evaluating the size of the NPs dispersion after lyophilization and resuspension (Figure S4).¹⁷

Antibacterial activity of AgNPs against Gram-negative and Gram-positive bacteria

The antimicrobial properties of AgNPs have demonstrated their broad spectra of activity against Gram-positive and Gram-negative bacteria, fungi, and viruses.^{18–21} To assess the properties of our AgNPs in comparison to other antibacterial products, we chose Gram-negative bacteria *E. coli* MG1665 and Gram-positive bacteria *S. aureus*.

About 2×10^5 cells were incubated with different concentrations of antibacterial agents for 2 h. After the incubation

time, 100 μ L of step-wise decimal dilutions up to 2×10^3 theoretical cells number were plated on a solid growth medium and incubated overnight at 37°C.

The number of CFU grown on the plate was counted and the results expressed as difference in viability (logarithmic-fold reduction) in comparison to the untreated control. Differently from conventional approaches that normally utilize the broth microdilution method to determine the minimal inhibitory concentration (MIC) for the determination of antimicrobial activity,²² the tests reported in this work were performed in plates because AgNPs were not stable in bacteria medium (Figure S5). In addition, we assessed the capability of AgNPs to kill the bacteria cells rather than evaluating the bacteriostatic power (MIC) of our preparations.

First, to evaluate AgNPs activity, different concentrations (from 10^{+2} to 10^{-2} ng/mL silver content) were tested (data not shown). To compare the antibacterial efficacy of AgNPs with that of the commercial products, concentrations showing a threefold logarithmic decrease in cells number were selected.

Among commercially available NPs, we chose citrate-stabilized AgNPs synthesized in an aqueous solution with slightly higher average size (40 nm AgNPs from Sigma, SNPs) and CSNPs commercially available for medical uses with a diameter of a few nanometers (Figure S6). Their chemico-physical characteristics are summarized in Table 1. Moreover, we performed the antimicrobial experiments with an equal concentration of silver ions in the form of AgNO₃ salt. The NP concentrations were evaluated by ICP-OES analysis measuring the amount of Ag ions in each sample.

The AgNPs prepared in the present work showed a superior antibacterial efficacy compared with the above-mentioned commercial AgNPs (Figures 2 and 3).

Although the discussion in literature on the possible impact of NPs in the administration of silver is controversial, our results suggest a strong improvement in antibacterial activity using AgNPs compared to Ag⁺ salts (Figures 2 and 3). A detailed elucidation of the antibacterial effect of Ag⁺ ions is still lacking; however, we can speculate on two possible reasons supporting such a strong difference. First, although the Ag⁺ ions uptake is normally controlled by the action of cationic pumps, AgNPs are able to be taken up by the cells exploiting nonspecific interactions, thus resulting in higher Ag content inside and/or across the cell wall.^{23–25} A second plausible explanation resides in the favorable redox potential of Ag⁰ atoms on the NP surface, which is expected to trigger the generation of large amounts of free radicals leading to cytotoxic reactive oxygen species production.^{12,26–28}

The low activity of SNPs compared to AgNPs can be explained in terms of different size: indeed, as expected, the 40 nm-sized SNPs possess significantly lower surface area leading to both decreased interactions with bacteria and reduced release of Ag⁺ ions.²⁹ We found high data variability using CSNPs. This is consistent with a low colloidal stability of the suspension attributable to a strongly elevated surface tension. Indeed, TEM analysis revealed that CSNPs were very small, that is, in the sub-5 nm range (Figure S6). Such a small size combined with their bare surface claimed in

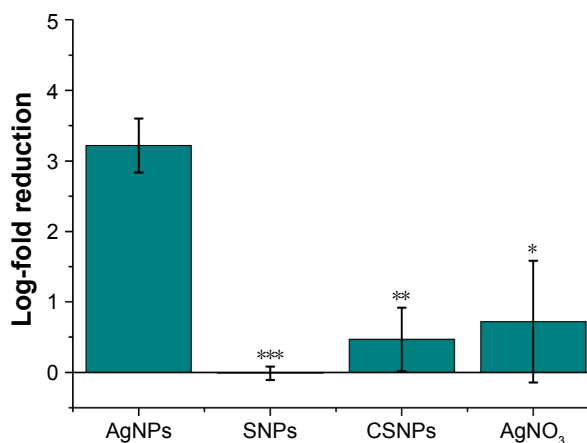


Figure 2 Antibacterial activity of AgNPs (0.54 ng/mL silver content) in *E. coli* MG1665 after 2 h. Samples show statistical differences from AgNPs samples at *** $p < 0.0005$; ** $p < 0.01$; * $p < 0.05$.

Abbreviations: AgNO₃, silver nitrate; AgNPs, silver nanoparticles; CSNPs, colloidal silver nanoparticles; *E. coli*, *Escherichia coli*; SNPs, silver nanoparticles.

product label results in a significantly reduced colloidal stability in biological media. In addition, the behavior of such small colloids is expected to be more similar to that of Ag ions than to larger NPs. Although CSNPs data were difficult to be interpreted due to such a low reproducibility, they suggested lower efficacy of CSNPs compared to AgNPs.

A further advance of our AgNPs relates to their negative surface charge, -18.3 ± 1.9 mV as determined by zeta potential; most examples from the literature suggest that an essential condition to improve the antibacterial efficacy of AgNPs is to have a positive particle surface charge. In fact, the positive charge allows for more efficient electrostatic interaction with the negative charges of the bacterial cell wall.³⁰ This expected effect obviously contrasted with our experimental data and represents a further important advantage of our antibacterial NPs in terms of safety in mammalian cells and tissues, assuming that cationic NPs are more cytotoxic than those with neutral or negative surface charge.³¹ Altogether, these properties make AgNPs an optimal alternative to cosmetic and food preservatives. As expected, the NP concentration necessary to obtain the same amount of dead cells was 10-fold higher in the case of *S. aureus* (Figure 3) in comparison

Table 1 Nanoparticles characterization

	Effective diameter (TEM analysis, nm)	PDI	Hydrodynamic diameter (DLS analysis, nm)	PDI	ζ potential (mV)	λ _{max} ABS (nm)
CSNPs	1.89±0.67	0.126	177.3±12.16	0.244	-17.8±0.802	–
SNPs	37.8±5.3	0.020	72.56±0.647	0.308	-36.6±0.643	417
AgNPs	19.3±5.6	0.084	42.3±4.7	0.203	-18.3±1.9	405

Abbreviations: AgNPs, silver nanoparticles; CSNPs, colloidal silver nanoparticles; SNPs, silver nanoparticles; DLS, dynamic light scattering; TEM, transmission electron microscopy.

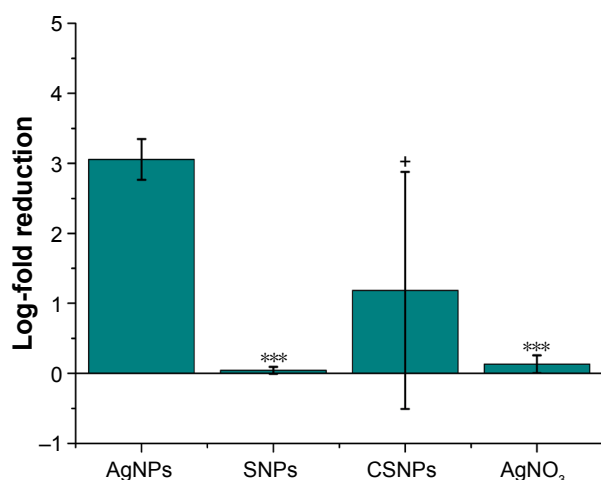


Figure 3 Antibacterial activity of AgNPs (5.4 ng/mL silver content) in *S. aureus* after 2 h. Samples show statistical differences from AgNPs samples at *** $P < 0.0005$; † $P < 0.2$.

Abbreviations: AgNO₃, silver nitrate; AgNPs, silver nanoparticles; CSNPs, colloidal silver nanoparticles; *S. aureus*, *Staphylococcus aureus*; SNPs, silver nanoparticles.

with *E. coli* (Figure 2) because of the presence of a thicker peptidoglycan layer in their cell wall.³²

Long-term antibacterial activity of AgNPs against Gram-negative bacteria

To evaluate the extent of the antimicrobial effect of AgNPs in comparison to CSNPs, the NPs were incubated with 2×10^5 *E. coli* cells for 2, 4, 6, and 24 h with a NP concentration lower than the above experiments in order to evaluate the inhibition effect over time. After the incubation time, 2×10^2 theoretical cell numbers of each sample were plated on a solid growth medium and incubated overnight at 37°C in duplicate. The number of CFU grown on the plate was counted and the results expressed in percentage of the untreated control. As expected, smaller AgNPs (CSNPs) exhibited a significantly higher release rate and final ion concentration compared to larger ones (AgNPs), resulting in an improved long-lasting effect of AgNPs (Figure 4).

In contrast, no prolonged effect was observed with Gram-positive bacteria due to a rapid and complete cell death already after 2 h.

AgNPs phase transfer in oil phase and antimicrobial properties maintenance

AgNPs are intended to be developed for pharmaceutical and cosmetic preparations administrable to humans and/or animals in liquid, semisolid, and solid pharmaceutical formulations on the need. As commercial formulations are based on water-soluble, fat-soluble ingredients and their

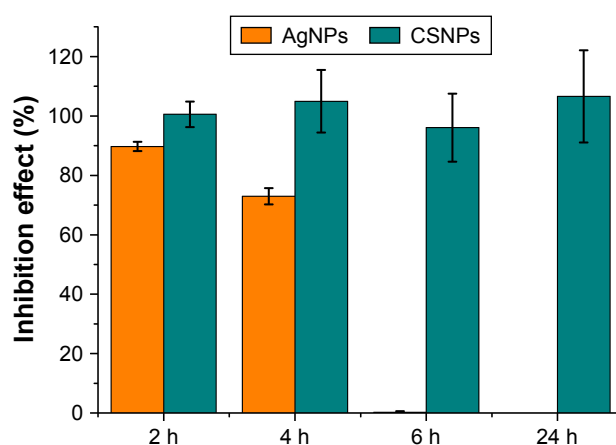


Figure 4 Long-term antibacterial activity of nanoparticles.

Note: AgNPs and CSNPs (0.054 ng/mL silver content) were incubated with *E. coli* MG1665 for 2, 4, 6, and 24 h.

Abbreviations: AgNPs, silver nanoparticles; CSNPs, colloidal silver nanoparticles; *E. coli*, *Escherichia coli*.

combinations, a phase transfer approach was developed to allow AgNPs to be suspended not only in aqueous solution but also in oily mixtures. Hence, a protocol to transfer AgNPs from the aqueous to the organic/oily phase was developed using octadecylamine as a surfactant added to the NPs suspension.³¹

The product was, thus, suspended in peanut oil or paraffin. The solution was homogeneous and stable (Figure S7). Differently from other methods, the advantage of the reported phase transfer protocol consists in its good transfer efficiency independently on the volume and the number of NPs.³¹

Tests analog to those described above were carried out using AgNPs soluble in oil phase. THF was chosen as an organic solvent to suspend AgNPs for its miscibility with a cell medium. The obtained data showed that the antibacterial efficacy of AgNPs was retained even when AgNPs were rendered suitable for the dispersion in oil. AgNPs exhibited the same concentration-dependent effect even if administered in an organic solution (Figure 5).

Cellular cytotoxicity of AgNPs

An MTT assay was performed to evaluate the toxicity of AgNPs on murine fibroblasts (3T3-L1 cell line). Tests were performed at the same AgNPs antibacterial concentrations selected in the previous experiments with Gram-positive and Gram-negative bacteria. The medium chosen for the cytotoxicity assay was DMEM, in accordance with the NP stability.

Water- or THF-suspended AgNPs were incubated at effective antimicrobial concentrations with 3T3-L1 murine

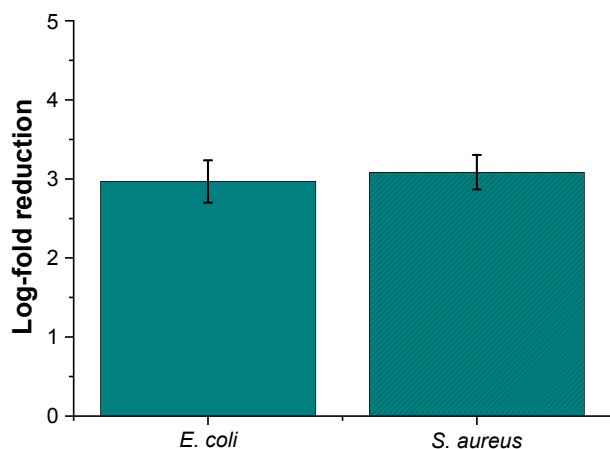


Figure 5 Antibacterial activity of AgNPs after the phase transfer, suspended in THF in *Escherichia coli* MGI665 (0.54 ng/mL silver content) and in *Staphylococcus aureus* (5.4 ng/mL silver content) after 2 h.

Abbreviations: AgNPs, silver nanoparticles; THF, tetrahydrofuran.

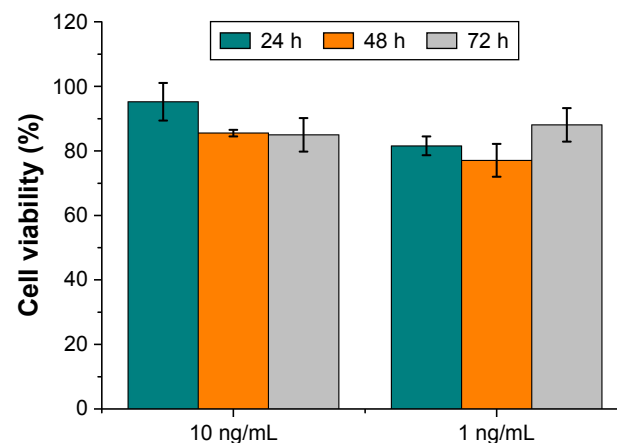


Figure 7 MTT assay performed incubating AgNPs suspended in THF (10 and 1 ng/mL silver content) with murine fibroblasts 3T3-L1.

Abbreviations: AgNPs, silver nanoparticles; MTT, 3-(4,5-dimethyl-2-thiazolyl)-2,5-diphenyl-2H-tetrazolium bromide; THF, tetrahydrofuran.

fibroblasts. ANOVA statistical analysis showed no significant differences ($P < 0.01$) except for both tested concentrations of AgNPs in THF at 48 h of incubation. Observed statistical differences do not have biological significance, as cell viability percentages were 85.5 and 77.1, respectively, for cells incubated with 10 and 1 ng/mL AgNPs, respectively. Indeed, AgNPs did not display significant cytotoxicity at the concentrations tested (Figures 6 and 7).

Scanning and transmission electron microscopy analyses

To evaluate the interaction between AgNPs and bacteria cells, both supernatants and cell pellets of *E. coli* and

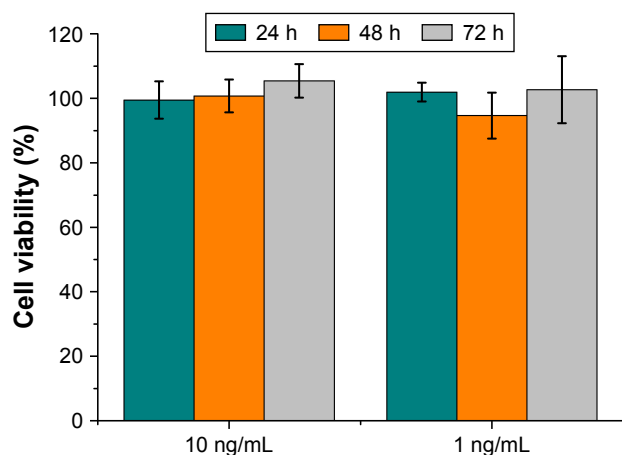


Figure 6 MTT assay performed incubating water-suspended AgNPs (10 and 1 ng/mL silver content) with murine fibroblasts 3T3-L1.

Abbreviations: AgNPs, silver nanoparticles; MTT, 3-(4,5-dimethyl-2-thiazolyl)-2,5-diphenyl-2H-tetrazolium bromide.

S. aureus obtained after 2-h incubation were visualized by TEM and SEM.

The untreated *E. coli* cells in the supernatant (Figure 8A and C) displayed a smooth and intact surface, whereas their interaction with AgNPs (Figure 8B) caused morphological changes in the cell wall visualized as membrane corrugations. In particular, electron-dense zones surrounding the cell wall appeared in the TEM image of treated bacteria (Figure 8D). This effect was attributable to a distress effect experienced by treated cells compared with untreated ones.¹²

Figure 9 shows the electron microscopy images of *S. aureus*. In the control sample (Figure 9A and C), the cells were spherical in shape and intact. The treated cells (Figure 9B and D) showed some blisters on the surface in contact with AgNPs. Furthermore, numerous debris around cells were observed: this phenomenon was probably caused by cells lysis, as confirmed by analysis of cell pellets by TEM (Figure 10). In addition, only treated samples displayed the presence of mesosome-like structures (Figure 10B) in the cytoplasm. No apparent cell lysis was recovered after incubation of AgNPs with cells. This observation suggested that other mechanisms should be involved in cell damage and death, for example, membrane permeability change.³³

Conclusion

In this study, uniform AgNPs having an average size of 20 nm and negative zeta potential with powerful antibacterial activity, suitable for pharmaceutical preparations administrable to humans and/or animals, are reported. Such NPs proved to

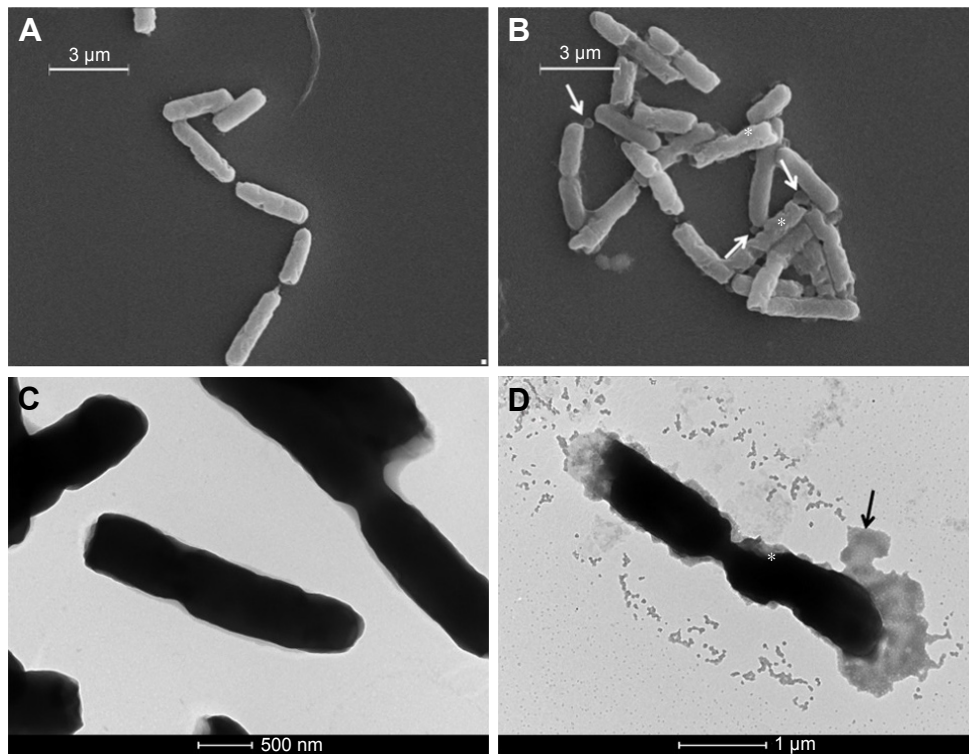


Figure 8 External morphology of untreated (**A, C**) and treated (**B, D**) *Escherichia coli* cells visualized with SEM (**A, B**) and TEM (**C, D**), respectively. Arrows indicate corrugations; * indicate membrane swelling (scale bars of panels **A, B** = 3 μm ; panel **C** = 500 nm; panel **D** = 1 μm).
Abbreviations: SEM, scanning electron microscopy; TEM, transmission electron microscopy.

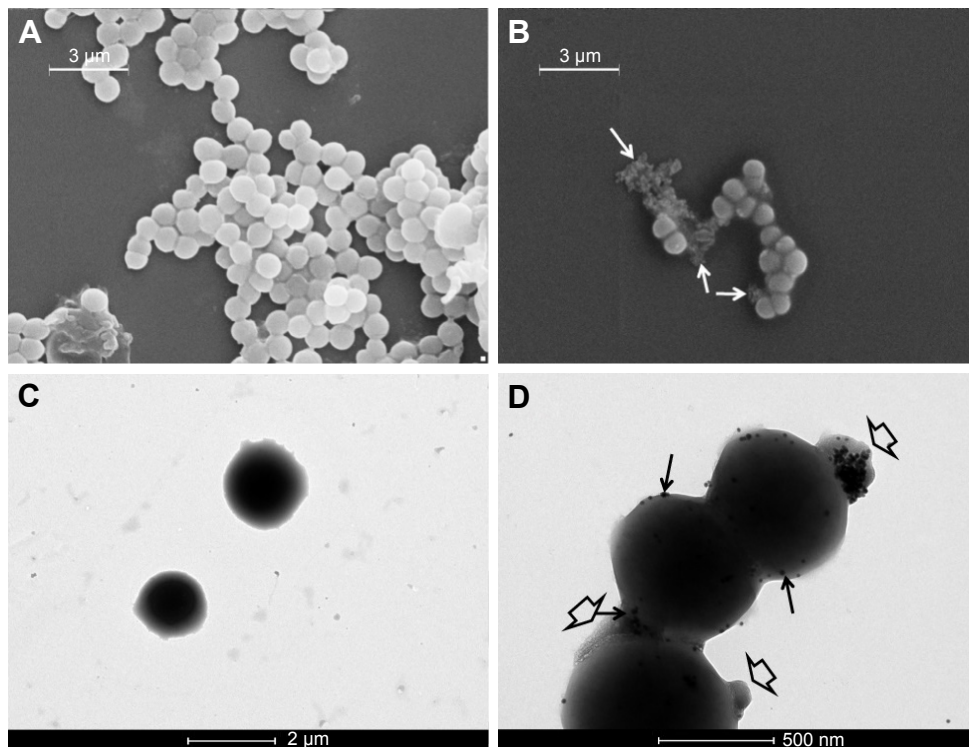


Figure 9 External morphology of untreated (**A, C**) and treated *Staphylococcus aureus* cells (**B, D**) visualized with SEM (**A, B**) and TEM (**C, D**), respectively. Big arrows indicate electron-dense AgNPs, whereas small arrows indicate swelling (scale bars of panels **A, B** = 3 μm ; panel **C** = 2 μm ; panel **D** = 500 nm).
Abbreviations: AgNPs, silver nanoparticles; SEM, scanning electron microscopy; TEM, transmission electron microscopy.

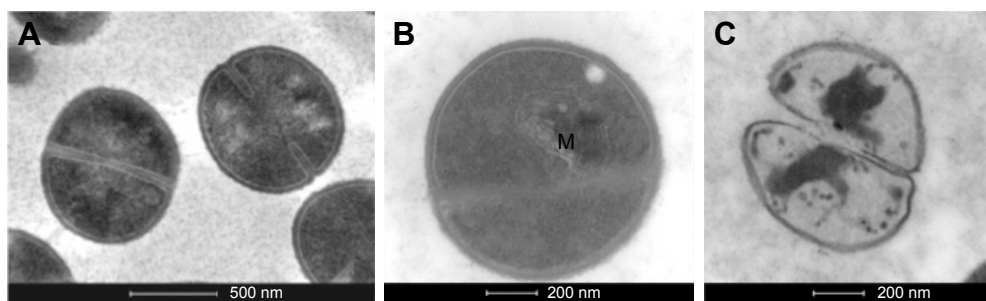


Figure 10 TEM images of untreated (A) and treated (B, C) *Staphylococcus aureus* cells from the pellet. In panel B, there is a cell that presents a mesosome-like particle (M). In panel C, a lysed cell is represented (scale bars of panel A =500 nm; panels B, C =200 nm).

Abbreviation: TEM, transmission electron microscopy.

be significantly more potent compared to currently available standard AgNPs at the same concentration of silver ions. In addition, AgNPs exhibited a durable antibacterial activity and retained their efficacy both in aqueous and nonaqueous media and after freeze-drying processes. Their easy and straightforward synthesis combined with very low toxicity for eukaryotic cells make them a promising alternative to conventional additives used in food, cosmetic, nutraceutical, and pharmaceutical formulations for the conservation and preservation of their microbiological integrity.

Taking advantage of the strong stability both in aqueous and oily solutions, AgNPs can be included in liquid, semi-solid, and solid formulations suitable to be administered by topical route, including intact or injured skin and mucosae, oral route, or applied on rigid or flexible surfaces that are in contact with the human or animal body.

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Disclosure

The authors report no conflicts of interest in this work.

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Supplementary materials

Table S1 Silver nanoparticles characterization

Sample	Effective diameter (TEM analysis, nm)	Hydrodynamic diameter (DLS analysis, nm)	ζ potential (mV)
1	19.3 \pm 5.6	41.4 \pm 2.0	21.6 \pm 1.8
2	17.7 \pm 6.9	41.5 \pm 1.8	23.5 \pm 2.9
3	20.9 \pm 4.7	42.3 \pm 4.7	18.3 \pm 1.9

Note: The results are expressed as mean \pm standard deviation of three measurements.

Abbreviations: DLS, dynamic light scattering; TEM, transmission electron microscopy.

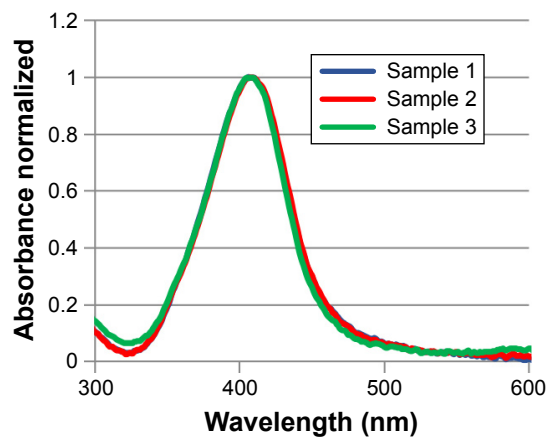


Figure S1 Plasmon band position of three samples of AgNPs by UV-vis absorption spectroscopy.

Abbreviations: AgNPs, silver nanoparticles; UV-vis, ultraviolet-visible.

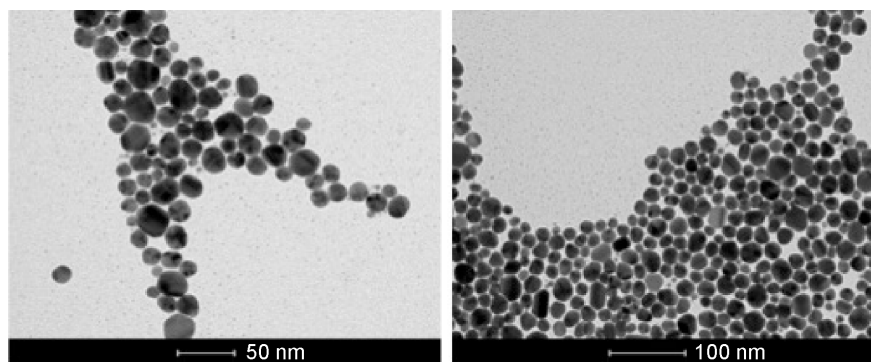


Figure S2 TEM images of AgNPs suspended in water, scale bar: 50 nm (left) and 100 nm (right).

Abbreviations: AgNPs, silver nanoparticles; TEM, transmission electron microscopy.

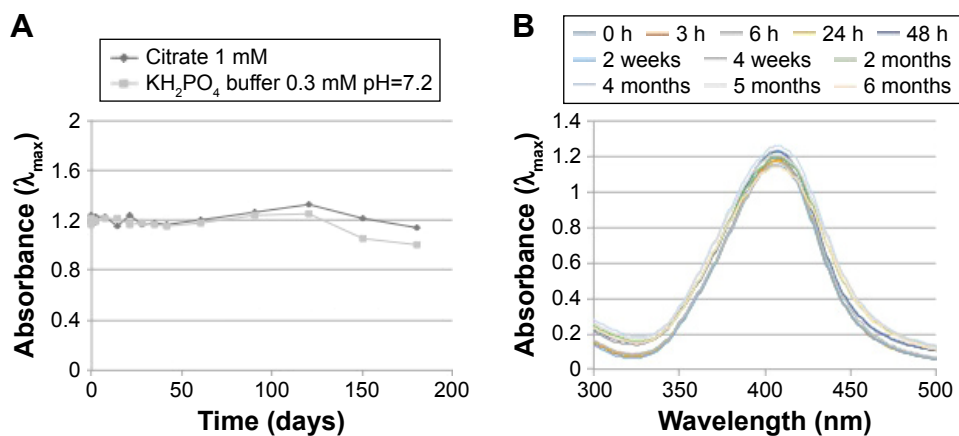


Figure S3 Long-term stability of AgNPs in citrate buffer 1 mM and phosphate buffer 0.3 mM (pH 7.2) (A) expressed as absorbance at λ_{\max} variations as function of time; plasmon band position of AgNPs by UV-vis absorption spectroscopy (B).

Abbreviations: AgNPs, silver nanoparticles; UV-vis, ultraviolet-visible.

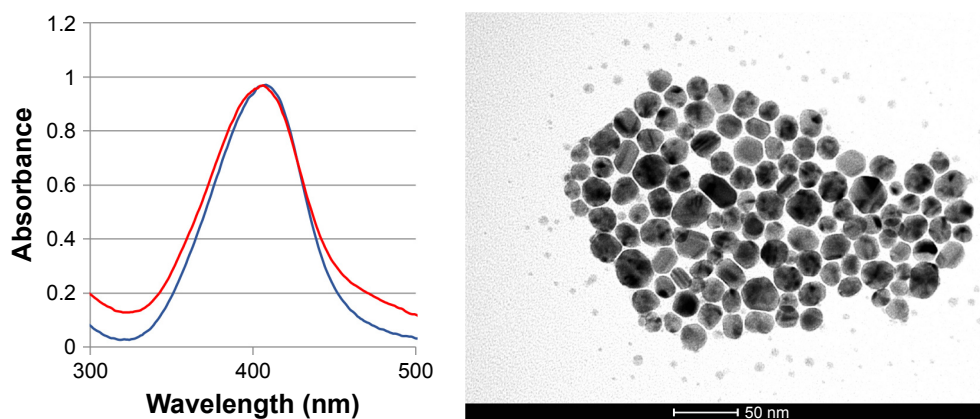


Figure S4 AgNPs stability evaluated by UV-vis spectroscopy in citrate 1 mM before (red line) and after (blue line) the lyophilization (left). TEM image of lyophilized AgNPs (right).

Abbreviations: AgNPs, silver nanoparticles; TEM, transmission electron microscopy; UV-vis, ultraviolet-visible.

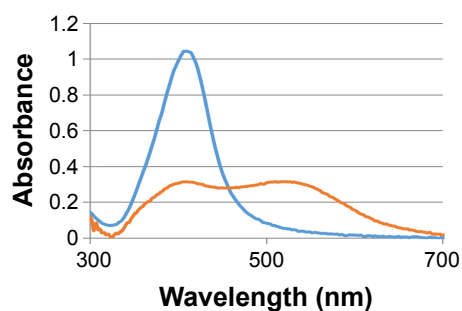


Figure S5 AgNPs suspended in bacteria nutrient broth (orange line) and in sodium citrate 1 mM solution (blue line). In the first case, as broader peak shows, nanoparticles aggregation occurs.

Abbreviation: AgNPs, silver nanoparticles.

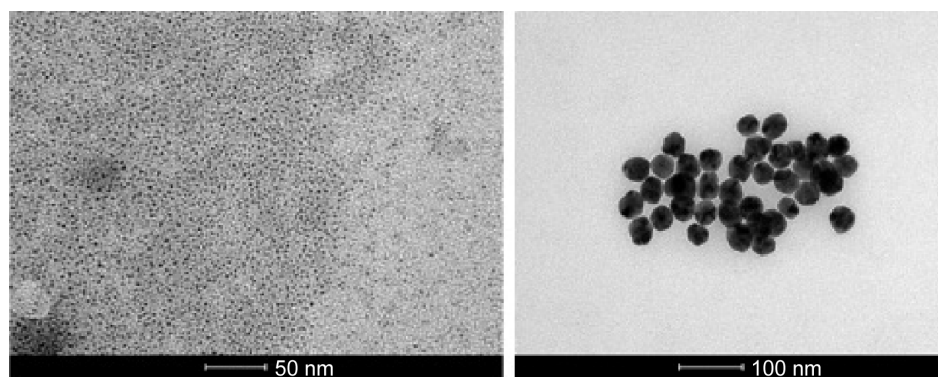


Figure S6 TEM images of CSNPs, scale bar: 50 nm (left) and SNPs scale bar: 100 nm (right).

Abbreviations: CSNPs, colloidal silver nanoparticles; SNPs, silver nanoparticles; TEM, transmission electron microscopy.

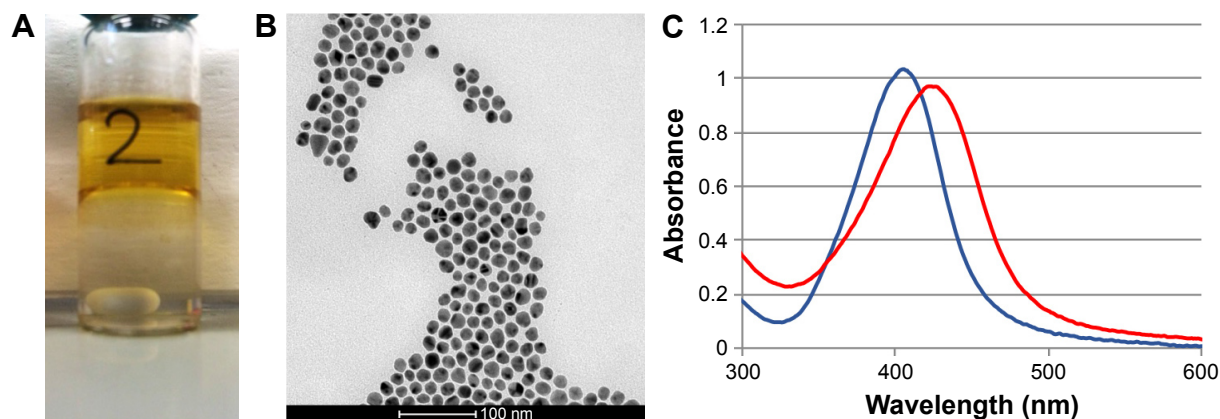


Figure S7 AgNPs in toluene characterization.

Notes: (A) AgNPs phase transfer in toluene; (B) TEM image of AgNPs in toluene; (C) AgNPs stability in water (blue line) after the phase transfer in toluene (red line). Differences of UV-vis spectra may be due to surface binding of octadecylamine and increasing in the refractive index from water (refractive index, $n=1.33$) to toluene ($n=1.50$).

Abbreviations: AgNPs, silver nanoparticles; TEM, transmission electron microscopy; UV-vis, ultraviolet-visible.

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